Chemistry 141 Name

Dr. Cary Willard

Exam 1b September 24, 2009

 Multiple Choice (30 points)

 Page 3 (10 points)

 Page 4 (18 points)

 Page 5 (15 points)

 Page 6 (18 points)

 Page 7 (16 points)

 Page 8 (15 points)

 Total (122 points)

 Percent (100 %)

All work must be shown to receive credit. Give all answers to the correct number of significant figures

Avogadros number = 6.022 x 1023 /mol

4 quarts = 1 gallon

36 in = 1 yard

Grossmont College

Periodic Table

|  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
|  IA |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | VIIA | NOBLE GASES |
| 1**H**1.008 | IIA |  |  |  |  |  |  |  |  |  |  | IIIA | IVA | VA | VIA | 1**H**1.008 | 2**He**4.002 |
| 3**Li**6.941 | 4**Be**9.012 |  |  |  |  |  |  |  |  |  |  | 5**B**10.81 | 6**C**12.01 | 7**N**14.01 | 8**O**16.00 | 9**F**19.00 | 10**Ne**20.18 |
| 11**Na**23.00 | 12**Mg**24.30 | IIIB | IVB | VB | VIB | VIIB |  VIII VIII VIII | IB | IIB | 13**Al**27.00 | 14**Si**28.09 | 15**P**30.97 | 16**S**32.06 | 17**Cl**35.45 | 18**Ar**39.95 |
| 19**K**39.10 | 20**Ca**40.08 | 21**Sc**44.96 | 22**Ti**47.90 | 23**V**50.94 | 24**Cr**52.00 | 25**Mn**54.94 | 26**Fe**55.85 | 27**Co**58.93 | 28**Ni**58.70 | 29**Cu**63.55 | 30**Zn**65.38 | 31**Ga**69.72 | 32**Ge**72.59 | 33**As**74.92 | 34**Se**78.96 | 35**Br**79.90 | 36**Kr**83.80 |
| 37**Rb**85.47 | 38**Sr**87.62 | 39**Y**88.91 | 40**Zr**91.22 | 41**Nb**92.91 | 42**Mo**95.94 | 43**Tc**(99) | 44**Ru**101.1 | 45**Rh**102.9 | 46**Pd**106.4 | 47**Ag**107.9 | 48**Cd**112.4 | 49**In**114.8 | 50**Sn**118.7 | 51**Sb**121.8 | 52**Te**127.6 | 53**I**126.9 | 54**Xe**131.3 |
| 55**Cs**132.9 | 56**Ba**137.3 | 57**La**138.9 | 72**Hf**178.5 | 73**Ta**180.9 | 74**W**183.9 | 75**Re**186.2 | 76**Os**190.2 | 77**Ir**192.2 | 78**Pt**195.1 | 79**Au**197.0 | 80**Hg**200.6 | 81**Tl**204.4 | 82**Pb**207.2 | 83**Bi**209.0 | 84**Po**(209) | 85**At**(210) | 86**Rn**(222) |
| 87**Fr**(223) | 88**Ra**226.0 | 89**Ac**227.0 | 104**Rf**(261) | 105**Db**(262) | 106**Sg**(263) | 107**Bh**(262) | 108**Hs**(265) | 109**Mt**(266) | 110**??**(269) |  |  |  |  |  |  |  |  |

|  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| 58**Ce**140.1 | 59**Pr**140.9 | 60**Nd**144.2 | 61**Pm**(147) | 62**Sm**150.4 | 63**Eu**152.0 | 64**Gd**157.3 | 65**Tb**158.9 | 66**Dy**162.5 | 67**Ho**164.9 | 68**Er**167.3 | 69**Tm**168.9 | 70**Yb**173.0 | 71**Lu**175.0 |
| 90**Th**232.0 | 91**Pa**231.0 | 92**U**238.0 | 93**Np**(237) | 94**Pu**(244) | 95**Am**(243) | 96**Cm**(247) | 97**Bk**(247) | 98**Cf**(251) | 99**Es**(252) | 100**Fm**(257) | 101**Md**(258) | 102**No**(259) | 103**Lr**(260) |

Lanthanide series

Actinide series

Part I – Multiple Choice (30 points)

1. Which one of the following statements about temperature scales is **false**?
	1. The Celsius degree is smaller than the Fahrenheit degree.
	2. The freezing point of water on the Celsius scale is 0 degrees.
	3. The boiling point of water on the Fahrenheit scale is 212 degrees.
	4. All temperatures on the Kelvin scale are positive numbers.
2. Which of the following statements does not describe a **physical** property of chlorine?
	1. The freezing point of chlorine is -101¶C.
	2. The color of chorine gas is green.
	3. The density of chlorine gas at standard temperature and pressure is 3.17 g/L.
	4. Chlorine combines with sodium to form table salt.



1. If hitting the bulls eye is the desired result, Figure (a) represents
	1. Good accuracy and poor precision
	2. Poor accuracy and poor precision.
	3. Poor accuracy and good precision
	4. Good accuracy and good precision
2. According to history, the concept that all matter is composed of atoms was first proposed by
	1. Dalton, and widely accepted within a few decades.
	2. Dalton, but not widely accepted until the work of Mendeleev.
	3. Dalton, but not widely accepted until the work of Einstein.
	4. the Greek philosopher Democritus, but not widely accepted until modern times.
3. Which are isotopes? An atom that has an atomic number of 34 and a mass number of 76 is an isotope of an atom that has
	1. an atomic number of 34 and a mass number of 80.
	2. 42 neutrons and 34 protons.
	3. an atomic number of 32 and a mass number of 76.
	4. 42 protons and 34 neutrons.
4. To the correct number of significant figures, what is the temperature reading on the following Celsius thermometer?
	1. 15oC
	2. 15.67 oC
	3. 16 oC
	4. 15.6 oC
5. Carbon dioxide is an example of
	1. an element.
	2. a compound.
	3. a heterogeneous mixture.
	4. a homogeneous mixture.
6. Which of the species below has 28 protons and 26 electrons?
	1. $Ni^{2+}$
	2. $$
	3. $Fe^{2+}$
	4. $$
7. The gas Freon-11, CCl3F, contains
	1. CCl3F molecules.
	2. C4+, Cl3-, and F- ions.
	3. C4+ and Cl3F4- ions.
	4. C4+, Cl-, and F- ions.
8. Which one of the following statements about balanced equations is **false**? In a balanced reaction
	1. mass must be conserved.
	2. net charge must be balanced on both sides of the reaction arrow.
	3. molecules must be balanced on both sides of the reaction arrow.
	4. atoms must be balanced on both sides of the reaction arrow.
9. Which statement about diluted solutions is **false**? When a solution is diluted
	1. the number of moles of solute remains unchanged.
	2. the number of moles of solvent remains unchanged.
	3. the concentration of the solution decreases.
	4. the molarity of the solution decreases.
10. Which statement about elemental analysis by combustion is **not** correct?
	1. Hydrogen is determined from the amount of H2O formed.
	2. Carbon is determined from the amount of CO2 formed.
	3. Oxygen is determined from the amount of H2O formed.
	4. Only carbon and hydrogen can be determined directly from CO2 and H2O.
11. What reagent could be used to separate Br- from NO3- when added to an aqueous solution containing both?
	1. NaI (*aq)*
	2. AgNO3 (*aq)*
	3. Ba(OH)2 (*aq)*
	4. CuSO4 (*aq)*
12. What is the oxidation number of the chromium atom in K2Cr2O7?
	1. -2
	2. +2
	3. +6
	4. +7
13. Which species functions as the reducing agent in the following reduction-oxidation reaction:

2 P(*s*) + 3 Br2(*l*) 🡪 2 PBr3(*l*)?

* 1. P(*s*)
	2. Br-(*aq*)
	3. Br2(*l*)
	4. P3+(*aq*)

Part 2 - Problems

1. (5 points) Give the IUPAC name for the following compounds
	1. Fe2(SO4)3
	2. (NH4)3PO4
	3. Br3O8
	4. Zn(MnO4)2
	5. NaClO4
2. (5 points) Write the correct formula for each of the following compounds
	1. potassium oxalate
	2. Nitrogen monoxide
	3. Cadmium borate
	4. cupric hydroxide
	5. carbonic acid
3. (6 points) The Honda Insight, a hybrid electric vehicle, has an EPA gas mileage rating of 43 mi/gal on the highway. How many kilometers can the Insight travel on the amount of gasoline that would fit in a soda pop can? The volume of a soda pop can is 355 mL.
4. (6 points) Nitrogen fixation in the root nodules of peas and other legumes occurs with a reaction involving a molybdenum containing enzyme named nitrogenase. This enzyme contains two Mo atoms per molecule and is 0.0635% Mo by mass. What is the molar mass of the enzyme?
5. (6 points) If 65.00 g of a mixture of 25.00% CaCl2 and 75.00% inert material are dissolved in 1.000 liter of water, how many mL of 0.500 M silver nitrate solution will be required to precipitate all of the chloride? The reaction equation is:

2 AgNO3 + CaCl2 🡪 Ca(NO3)2 + 2 AgCl

1. (15 points) Quinine is often used as an anti-malarial drug. Its molecular formula is C20H24O2N2. Answer the following questions regarding quinine.
	1. Calculate the molar mass of quinine.
	2. Calculate the mass of quinine that contains 5.923 x 1025 atoms of carbon.
	3. Calculate the number of moles of hydrogen in 2.66 moles of quinine.
	4. Calculate the number of molecules of quinine that contains 436 atoms of oxygen.
	5. Calculate the mass in grams of one molecule of quinine.
2. (9 points) Complete the following precipitation reaction with balanced molecular, total ionic, and net ionic equations.
	1. Cr2(CO3)3*(s)* + HNO3*(aq)*
	2. Balanced total ionic equation
	3. Balanced net ionic equation
3. (9 points) Balance the following reaction in acid

Cr2O6-2 + Cu+1 🡪 Cr+3 + Cu+2

Oxidation half reaction

Reduction half reaction

Overall balanced equation

1. (8 points) A solution of nitric acid is prepared by diluting 21.6 mL of a 12.0 M solution of nitric acid to 25.00 L with water.
	1. What is the final concentration of nitric acid in the dilute solution?
	2. What is the pH of the final solution?
2. (8 points) A noncarbonated soft drink contains an unknown amount of citric acid H3C6H5O2. If 100.0 mL of the soft drink requires 45.82 mL of 0.1020 M NaOH to neutralize the citric acid completely, what is the molarity of citric acid in the soft drink?
3. (15 points) You mix 527.0 mL of 0.2754 M sodium phosphate with 300.0 mL of 0.6684 M vanadium(II) chloride. Write the reaction and determine the number of grams of vanadium(III) carbonate produced, and the final concentration of all ions in the solution.

Balanced chemical equation (Check with me before you go on to be sure this is correct.)

|  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- |
|  | x =  |  | x =  |  |  |  |  |
|  |  | + |  | 🡪 |  | + |  |
| I |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |
| E |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |

Moles V3(PO4)2 produced Mass V3(PO4)2 produced

Moles Na+1 = [Na+1] =

Moles PO4-3 = [PO4-3] =

Moles V+2 = [V+2] =

Moles Cl-1 = [Cl-1] =